Chemistry 115 Name

Dr. Cary Willard

Quiz 4a (20 points) March 2, 2010

Must show all work to receive credit. Use proper significant figures.

Avogadro’s number – 6.022 x 1023 particles/mol

1. (4 points) How many atoms of silver are in a 0.578 mol sample of silver?
2. (4 points) A chunk of sugar contains 6.87 x 1026 molecules. How many moles of sugar are there in the sample?
3. (4 points) How many moles of chromium are in a 72.4 g sample of chromium?
4. (4 points) How many atoms of sodium are there in a 6.00 g sample of sodium?
5. (4 points) What is the molar mass of barium phosphite, Ba3(PO3)2?

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Quiz 4b (20 points) March 2, 2010

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Avogadro’s number – 6.022 x 1023 particles/mol

1. (4 points) How many atoms of silver are in a 0.384 mol sample of silver?
2. (4 points) A chunk of sugar contains 3.49 x 1026 molecules. How many moles of sugar are there in the sample?
3. (4 points) How many moles of chromium are in a 68.4 g sample of chromium?
4. (4 points) How many atoms of sodium are there in a 8.00 g sample of sodium?
5. (4 points) What is the molar mass of strontium phosphite, Sr3(PO3)2?